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1. If water sample are taken from sea, rivers or lake, they will be found to contain hydrogen and oxygen in the approximate ratio of 1 : 8. This indicates the law of :

- (a) Multiple proportion
- (b) Definite proportion**
- (c) Reciprocal proportions
- (d) None of these

2. Hydrogen and oxygen combine to form H_2O_2 and H_2O containing 5.93% and 11.2% hydrogen respectively. The data illustrates :

- (a) Law of conservation of mass
- (b) Law of constant proportion**
- (c) Law of reciprocal proportion
- (d) Law of multiple proportion

4. All the substances listed below are fertilizers that contribute nitrogen to the soil. Which of these is the richest source of nitrogen on a mass percentage basis?

- (a) Urea, $(\text{NH}_2)_2\text{CO}$
- (b) Ammonium nitrate, NH_4NO_3
- (c) Nitric oxide, NO
- (d) Ammonia, NH_3**

7. 20 g of an ideal gas contains only atoms of S and O occupies 5.6 L at NPT. What is the mol. wt. of gas ?

- (a) 64
- (b) 80**
- (c) 96
- (d) None of these

8. A sample of ammonium phosphate, $(\text{NH}_4)_3\text{PO}_4$, contains 6 moles of hydrogen atoms. The number of moles of oxygen atoms in the sample is:

- (a) 1
- (b) 2
- (c) 4**
- (d) 6

9. Total number of moles oxygen atoms in 3 litre O_3 (g) at $27^\circ C$ and 8.21 atm are:

- (a) 3**
- (b) 1
- (c) 1
- (d) None of these

10. 3.01110^{22} atoms of an element weight 1.15 gm. The atomic mass of the element is:

- (a) 10
- (b) 2.
- (c) 35.5
- (d) 23**